

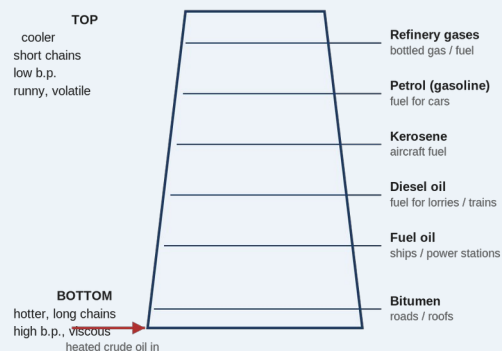
4.2 Crude Oil

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1. KEY VOCABULARY

TERM	MEANING
Crude oil	A finite mixture of hydrocarbons, mostly alkanes.
Fraction	A group of hydrocarbons with similar boiling points.
Fractional distillation	Separating crude oil into fractions by boiling point.
Viscosity	How thick / how easily a liquid flows.
Volatility	How easily a substance turns into a vapour.
Cracking	Breaking long-chain hydrocarbons into shorter, more useful ones.

2. FRACTIONAL DISTILLATION



3. HOW IT WORKS

Crude oil is heated and the vapours rise up the column. **The column is hot at the bottom, cool at the top.** Each fraction condenses where the temperature matches its boiling point. Short chains rise highest; long chains condense low down.

4. TRENDS DOWN THE COLUMN

Going DOWN (longer chains):

- higher boiling point
- more viscous (thicker)
- less volatile, less flammable
- darker in colour

5. CRACKING

Long-chain hydrocarbons are less useful. Cracking breaks them into shorter alkanes and alkenes.

long alkane → shorter alkane + alkene

Uses heat and a catalyst. It also supplies alkenes for making polymers.

6. THE WHY

Why fractional distillation works: hydrocarbons of different chain length have different boiling points, so they condense at different heights.

Why we crack oil: demand for short-chain fuels is higher than supply — cracking converts the surplus long chains.

7. COMMON EXAM MISTAKES

- ✗ "Crude oil is a compound."
- ✓ It is a MIXTURE of hydrocarbons.
- ✗ "Long chains are more volatile."
- ✓ Long chains are LESS volatile, more viscous.
- ✗ "Cracking joins molecules together."
- ✓ Cracking BREAKS long molecules into shorter ones.

8. SELF-CHECK · cover & quiz

Can you...

1. Explain what crude oil is?
2. Describe how fractional distillation separates it?
3. Name three fractions and a use for each?
4. State the trends in properties down the column?
5. Explain what cracking is and why it is done?
6. Write a general equation for cracking?